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1 Chemical
reactions 1.1
Rate of reaction
and dependence
on temperature
We will once
again look at
the formation of
ammonia (NH_3)
from nitrogen
and hydrogen
(see section
Chemical

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equilibrium of
the
thermodynamics
chapter). This
reaction follows
the equation: $N_2 + 3H_2 \rightleftharpoons 2NH_3$
(1) $H^0 = 92 \text{ kJ}$
 mol^{-1} $S^0 = 192 \text{ J}$
 $\text{mol}^{-1} \text{K}^{-1}$

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the sum of the

internal energy

U , the kinetic

energy K , and

the potential

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energy E_P The kinetic and the potential energy are negligible in many chemical reaction engineering applications, so Equation 15 becomes $Q - V E = Q - V (U + K + E_P) = Q - V U$ (16) we know that U is a function of the

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enthalpy,

pressure, and

volume, so $Q = V$

$U = Q - V (H - PV) = A$

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A first order
reaction is to
be treated in a
series of two

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Chemical reactors.

The total volume
of the two
reactors is

minimum, when

the reactors are

(A) Equal in

size (B) Of

different sizes

(C) Of such size

that the ratio

of their volumes

is < 5 (D) None

of these.

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Answer: Option A

• 23. Half life
period of a
first order

irreversible
reaction A \rightarrow B
is (A) $k/2$

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Test: Question

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case of staged

packed bed

reactors

carrying out

exothermic

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reaction, use

(A) High recycle
for pure gas (B)

Plug flow for

dilute liquid

requiring no

large preheating

of feed ...

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Prof. Milorad

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Thermodynamics,

Rates, Kinetics,

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graphical
procedures, and
frequent
comparison of
capabilities of
major reactor
types. This
approach helps
students develop
a strong
intuitive sense
for good design.

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engineering is a

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reactors.

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Frequently the term relates specifically to catalytic reaction systems where either a homogeneous or heterogeneous catalyst is present in the reactor.

Sometimes a reactor per se is not present

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by itself, but rather is integrated into a process, for example in reactive separations vessels, retorts, certain fuel cells, and photocatalytic surfaces. The issue of solvent effect

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chemistry can be
overcome by the
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while engineers
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definition of the

activity

of species i is:

fugacity at the
standard state,
that is, 1 atm
for gases and

thus. 3. $K = \frac{[N_2]^{1/2} [O_2]^{1/2}}{[NO]}$

$f^{3/2}(T, P) N, H,$

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Use of the
Lewis and
Randall rule
gives:

$$\left[\frac{f_{NH_3}}{f_{NH_3}^0} \right] = \left[\frac{f_{NH_3}}{f_{NH_3}^0} \right]_{I \text{ atm.}}$$

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Gavhane. ammonia

(NH₃) from

nitrogen and

hydrogen (see

section Chemical

equilibrium of

the

thermodynamics

chapter). This

reaction follows

the equation: N

$2 + 3H_2 \rightleftharpoons 2NH_3$

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 $\text{mol } SO = 192 \text{ J}$
 $\text{mol } K$

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