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Stoichiometry – Compositions Study  
Guide 1. Formulas and Subscripts

How many moles of hydrogen atoms  
are in 1 mole of each of the following

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compounds?  $H_2O$  2  $CH_3$   
 $OH$  4  $(NH_4)_2CO_3$  8

## 2. Empirical and Molecular Formulas

Give the empirical formula for each

compound:  $N_2H_2$            $NH$            $BH_3$   
          $BH_3$            $H_2O_2$            $OH$  ...

## Stoichiometry Compositions Study

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Stoichiometry. The atomic ratios in each compound are also the relative number of atomic mass units of its elements. The first example is nitrous oxide ( $N_2O$ ), as shown in Table 1. The relative masses were obtained by multiplying the atomic ratios and

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atomic masses. ... CliffsNotes study guides are written by real teachers and professors, so ...

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Those who doubt Jesus have sinned willfully, trampled under foot the Son of God, regarded as unclean the blood of Page 10/28

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Calculations In your textbook, read  
about mole-to-mole conversion. ...

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reaction represented by the equation  
 $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ , how

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many grams of hydrogen are  
produced if 120. g of Na and 80.0 g

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equation. For each statement below,  
write true or false. \_\_\_\_\_ true \_\_\_\_\_ 1.

The study of the quantitative

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relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.

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Stoichiometry Section 11.1 What is stoichiometry? 1. true 2. true 3. false 4. true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4, 4, 72.08 10. methanol and oxygen gas 11. carbon dioxide and water 12. 160.10 g 13. 160.10 g 14. They are equal. 15. A



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Mole ratio is a ratio between ...

## Answer

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